Chapter 09 WS

Name: _____

Names/Formulas of Molecular Compounds **Period:** _____ **Date:** _____

Directions

Name the Following Molecular Compounds

1) O_2 _____

2) H_2 _____

3) Cl_2 _____

4) F_2 _____

5) C_8H_{10} _____

6) SF_6 _____

7) P_4S_5 _____

8) SeF_6 _____

9) Si_2Br_6 _____

10) SCl_4 _____

11) PCl_5 _____

12) P_7I_9 _____

13) N_2O_4 _____

14) Cl_2O_7 _____

15) C_3O_4 _____

16) BF_3 _____

Directions

Write the formula for the following Molecular Compounds

1) Antimony tribromide _____

2) Hexaboron trisilicide _____

3) Tetrachlorine dioxide _____

4) Dinitrogen Trioxide _____

5) Octasilicon tetracarbide _____

6) Ammonia (nitrogen trihydride)_____

7) Phosphorous triiodide _____

8) Methane (carbon tetrahydride)_____

9) Nonasulfur heptabromide _____

10) Carbon tetrachloride _____

11) Trinitrogen Pentoxide _____

12) Diboron pentasilicide _____

13) Carbon monoxide _____

14) Carbon dioxide _____

Directions
Fill in the blanks.

1) A base is a compound that contains a _____ anion. This _____ anion is usually ionically bound to a metal cation and is named the exact same way as all ionic compounds with the _____ first, followed by the _____.
2) An acid is a compound that contains a _____ cation when placed in water. There are three basic types of acids. Acids can be a combination of the _____ cation with an anion that has a name that ends in either _____, _____, or _____. The name of the acid depends on the anion.

Directions
Write the formula for the following acids and bases.

3) Hydrochloric acid _____
4) Hydroiodic acid _____
5) Sulfuric acid _____
6) Nitric acid _____
7) Acetic acid _____
8) Nitrous acid _____
9) Chlorous acid _____
10) Sodium hydroxide _____
11) Calcium hydroxide _____
12) Vanadium (II) hydroxide _____

Directions
Write the name for the following acids and bases.

13) HBr _____
14) HF _____
15) H_3PO_4 _____
16) H_2CrO_4 _____
17) H_2CO_3 _____
18) HNO_2 _____
19) H_2SO_3 _____
20) KOH _____
21) $Pb(OH)_4$ _____
22) $Cr(OH)_3$ _____