

Electrochemistry

Standard Reduction Potentials (E^0)

Half-Reaction	V	Half-Reaction	V
$\text{Li}^+ + \text{e}^- \rightarrow \text{Li}$	-3.0401	$\text{Sn}^{+4} + 2\text{e}^- \rightarrow \text{Sn}^{+2}$	+0.151
$\text{Cs}^+ + \text{e}^- \rightarrow \text{Cs}$	-3.026	$\text{Cu}^{+2} + \text{e}^- \rightarrow \text{Cu}^+$	+0.153
$\text{K}^+ + \text{e}^- \rightarrow \text{K}$	-2.931	$\text{SO}_4^{-2} + 4\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2\text{SO}_3 + \text{H}_2\text{O}$	+0.172
$\text{Ba}^{+2} + 2\text{e}^- \rightarrow \text{Ba}$	-2.912	$\text{Bi}^{+3} + 3\text{e}^- \rightarrow \text{Bi}$	+0.308
$\text{Ca}^{+2} + 2\text{e}^- \rightarrow \text{Ca}$	-2.868	$\text{Cu}^{+2} + 2\text{e}^- \rightarrow \text{Cu}$	+0.3419
$\text{Na}^+ + \text{e}^- \rightarrow \text{Na}$	-2.71	$\text{O}_2 + 2\text{H}_2\text{O} + 4\text{e}^- \rightarrow 4\text{OH}^-$	+0.401
$\text{Mg}^{+2} + 2\text{e}^- \rightarrow \text{Mg}$	-2.372	$\text{Cu}^+ + \text{e}^- \rightarrow \text{Cu}$	+0.521
$\text{Ce}^{+3} + 3\text{e}^- \rightarrow \text{Ce}$	-2.336	$\text{I}_2 + 2\text{e}^- \rightarrow 2\text{I}^-$	+0.5355
$\text{H}_2 + 2\text{e}^- \rightarrow 2\text{H}^-$	-2.323	$\text{O}_2 + 2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2\text{O}_2$	+0.695
$\text{Nd}^{+3} + 3\text{e}^- \rightarrow \text{Nd}$	-2.1	$\text{Fe}^{+3} + \text{e}^- \rightarrow \text{Fe}^{+2}$	+0.771
$\text{Be}^{+2} + 2\text{e}^- \rightarrow \text{Be}$	-1.847	$\text{NO}_3^- + 2\text{H}^+ + \text{e}^- \rightarrow \text{NO}_2 + \text{H}_2\text{O}$	+0.775
$\text{Al}^{+3} + 3\text{e}^- \rightarrow \text{Al}$	-1.662	$\text{Hg}_2^{+2} + 2\text{e}^- \rightarrow 2\text{Hg}$	+0.7973
$\text{Mn}^{+2} + 2\text{e}^- \rightarrow \text{Mn}$	-1.185	$\text{Ag}^+ + \text{e}^- \rightarrow \text{Ag}$	+0.7996
$\text{Cr}^{+2} + 2\text{e}^- \rightarrow \text{Cr}$	-0.913	$\text{Hg}^{+2} + 2\text{e}^- \rightarrow \text{Hg}$	+0.851
$\text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{H}_2 + 2\text{OH}^-$	-0.8277	$2\text{Hg}^{+2} + 2\text{e}^- \rightarrow \text{Hg}_2^{+2}$	+0.920
$\text{Zn}^{+2} + 2\text{e}^- \rightarrow \text{Zn}$	-0.7618	$\text{Pd}^{+2} + 2\text{e}^- \rightarrow \text{Pd}$	+0.951
$\text{Cr}^{+3} + 3\text{e}^- \rightarrow \text{Cr}$	-0.744	$\text{NO}_3^- + 4\text{H}^+ + 3\text{e}^- \rightarrow \text{NO} + 2\text{H}_2\text{O}$	+0.957
$\text{Ga}^{+3} + 3\text{e}^- \rightarrow \text{Ga}$	-0.549	$\text{Br}_{2(l)} + 2\text{e}^- \rightarrow 2\text{Br}^-$	+1.066
$2\text{CO}_2 + 2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2\text{C}_2\text{O}_4$	-0.49	$\text{Ir}^{+3} + 3\text{e}^- \rightarrow \text{Ir}$	+1.156
$\text{S} + 2\text{e}^- \rightarrow \text{S}^{-2}$	-0.47627	$\text{Pt}^{+2} + 2\text{e}^- \rightarrow \text{Pt}$	+1.18
$\text{Fe}^{+2} + 2\text{e}^- \rightarrow \text{Fe}$	-0.447	$\text{O}_2 + 4\text{H}^+ + 4\text{e}^- \rightarrow 2\text{H}_2\text{O}$	+1.229
$\text{Cr}^{+3} + \text{e}^- \rightarrow \text{Cr}^{+2}$	-0.407	$\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$	+1.35827
$\text{Cd}^{+2} + 2\text{e}^- \rightarrow \text{Cd}$	-0.4030	$\text{Au}^{+3} + 2\text{e}^- \rightarrow \text{Au}^+$	+1.401
$\text{PbI}_2 + 2\text{e}^- \rightarrow \text{Pb} + 2\text{I}^-$	-0.365	$\text{Au}^{+3} + 3\text{e}^- \rightarrow \text{Au}$	+1.498
$\text{PbSO}_4 + 2\text{e}^- \rightarrow \text{Pb} + \text{SO}_4^{-2}$	-0.3588	$\text{MnO}_4^- + 8\text{H}^+ + 5\text{e}^- \rightarrow \text{Mn}^{+2} + 4\text{H}_2\text{O}$	+1.507
$\text{Co}^{+2} + 2\text{e}^- \rightarrow \text{Co}$	-0.28	$\text{Au}^+ + \text{e}^- \rightarrow \text{Au}$	+1.692
$\text{Ni}^{+2} + 2\text{e}^- \rightarrow \text{Ni}$	-0.257	$\text{H}_2\text{O}_2 + 2\text{H}^+ + 2\text{e}^- \rightarrow 2\text{H}_2\text{O}$	+1.776
$\text{Sn}^{+2} + 2\text{e}^- \rightarrow \text{Sn}$	-0.1375	$\text{Co}^{+3} + \text{e}^- \rightarrow \text{Co}^{+2}$	+1.92
$\text{Pb}^{+2} + 2\text{e}^- \rightarrow \text{Pb}$	-0.1262	$\text{S}_2\text{O}_8^{-2} + 2\text{e}^- \rightarrow 2\text{SO}_4^{-2}$	+2.010
$\text{Fe}^{+3} + 3\text{e}^- \rightarrow \text{Fe}$	-0.037	$\text{O}_3 + 2\text{H}^+ + 2\text{e}^- \rightarrow \text{O}_2 + 2\text{H}_2\text{O}$	+2.076
$2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$	0.00000	$\text{F}_2 + 2\text{e}^- \rightarrow 2\text{F}^-$	+2.866